

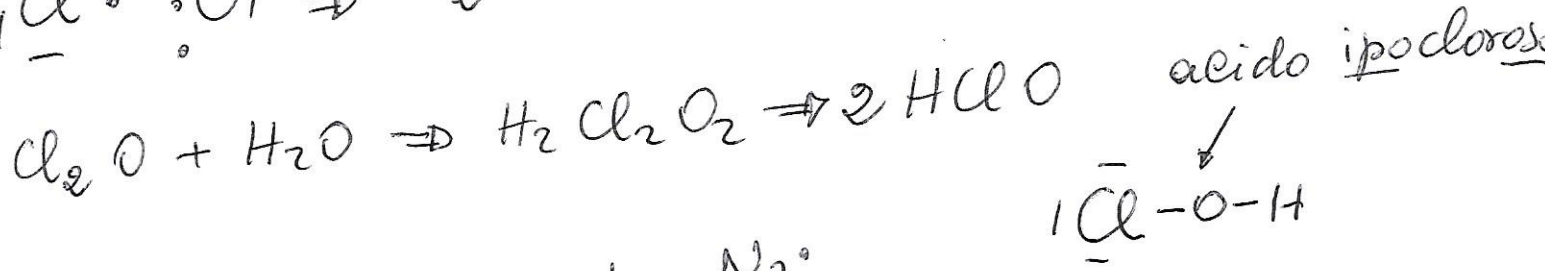
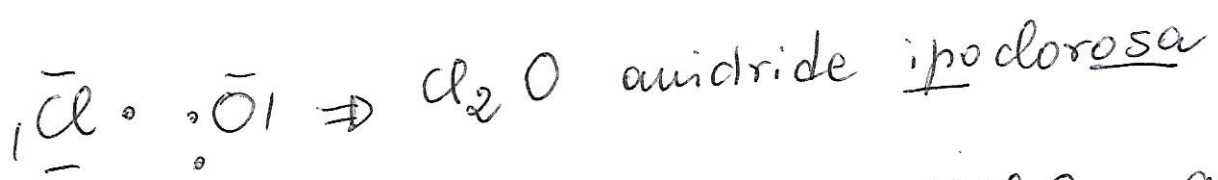
Catalogare il composto e scrivere la formula

ipoclorito di sodio

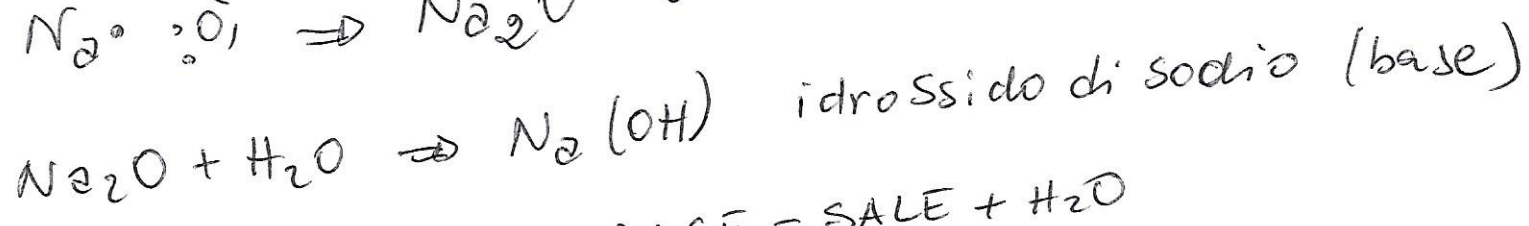
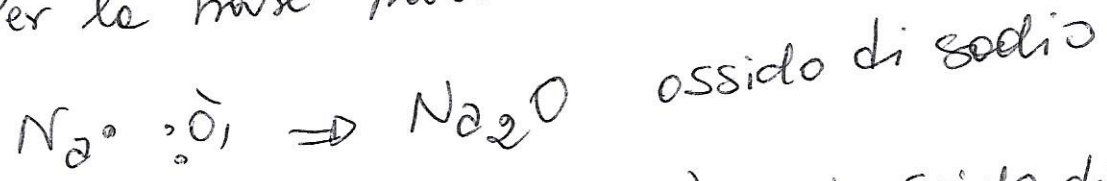
R. ITO indica un sale \Rightarrow devo trovare l'acido e la base (idrossido) che lo generano

IPO ITO indica che l'elemento da cui deriva l'acido ha più di 2 n. di ossidazione per il Cl si ha

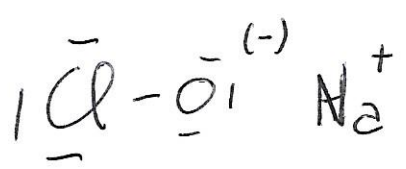
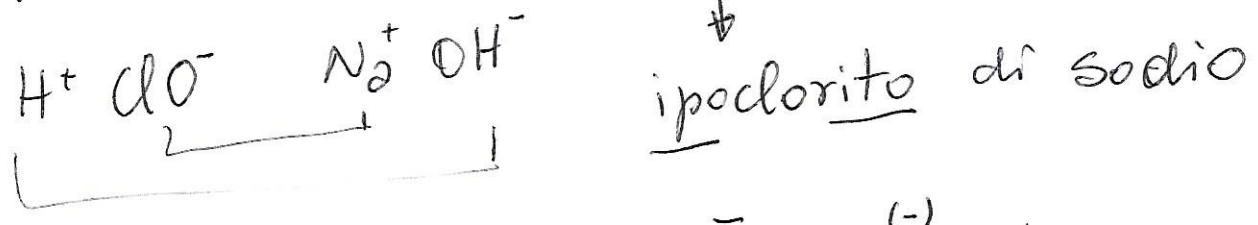
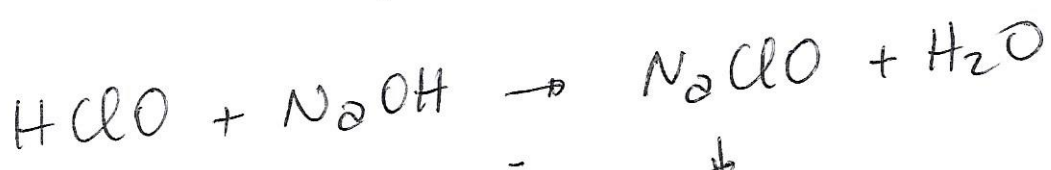
n_1	n_2	n_3	n_4	1	3	5	7
ipo	oso	ato	per	ipo			
oso			ato	oso	\Rightarrow devo riferirmi al n.o. 1		



* Per la base parto da Na^+



ACIDO + BASE = SALE + H_2O



2) Solfato di potassio

ATO ⇒ sale derivante da acido (ICO) + base
 d'elemento ha due n.o. prevalenti e devo considerare il massimo

\bar{S} : n.o. 6, 4 devo usare 6

$\bar{S}^{+6} : \bar{O}^{-2}$ ⇒ SO_3 anidride solforica

$SO_3 + H_2O \rightarrow H_2SO_4$ acido solforico

* da base si fa a partire da K

$K : \bar{O}$ K_2O ossido di potassio

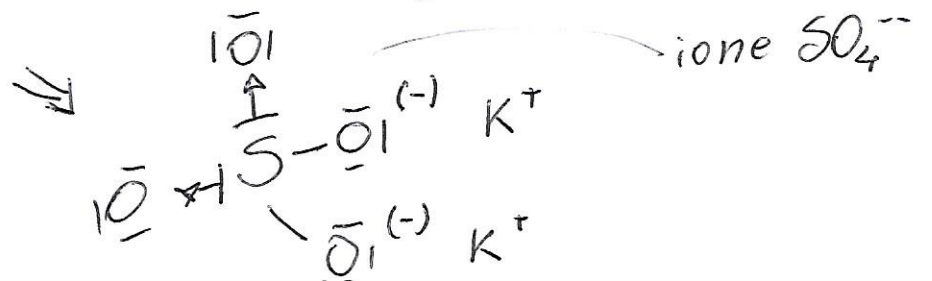
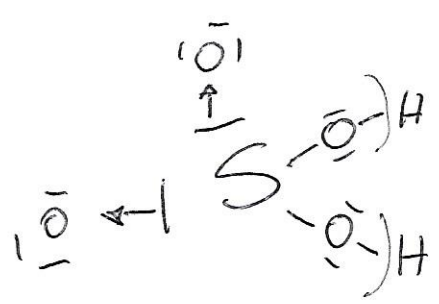
$K_2O + H_2O \rightarrow 2K(OH)$ idrossido di potassio

SALE = ACIDO + BASE

$H_2SO_4 + 2KOH \rightarrow K_2SO_4 + 2H_2O$

$2H^+ + [SO_4]^{-2} \rightarrow K^+(OH)^-$

↑
 servono due K^+ per bilanciare la doppia carica negativa di $[SO_4]^{--}$

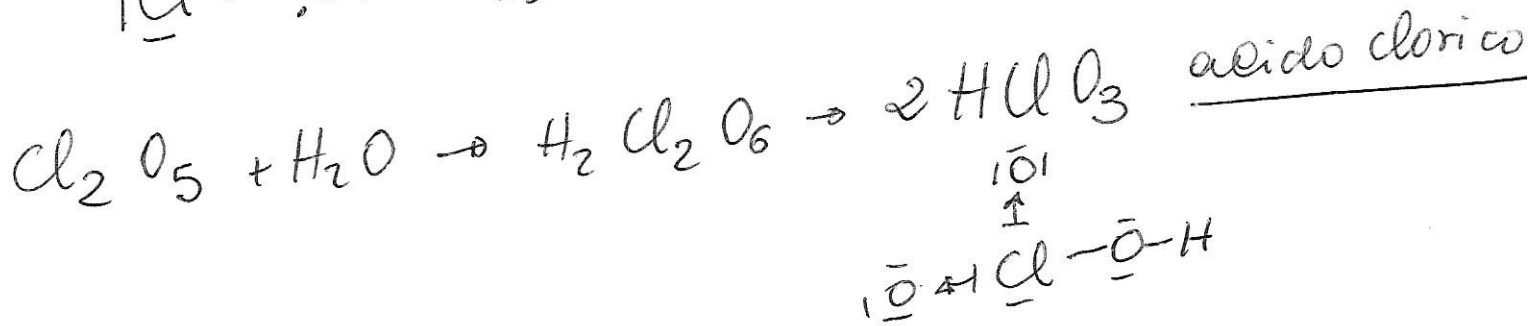
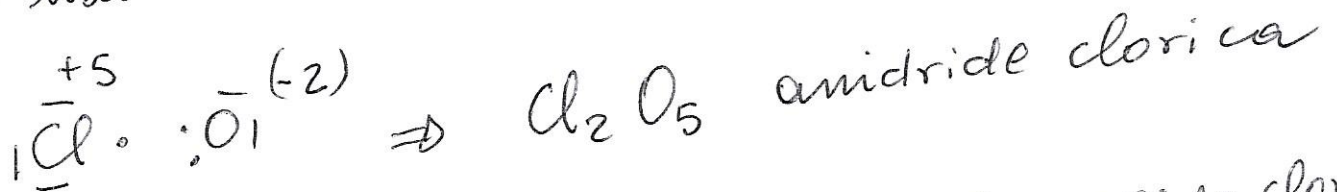


3) Clorato ferrico

ATO indica un sale proveniente da acido (ico)

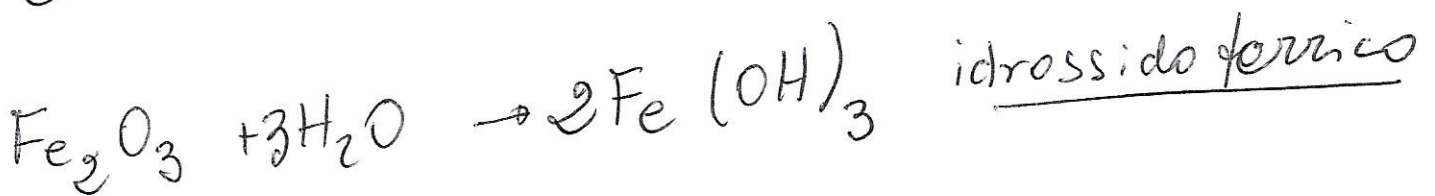
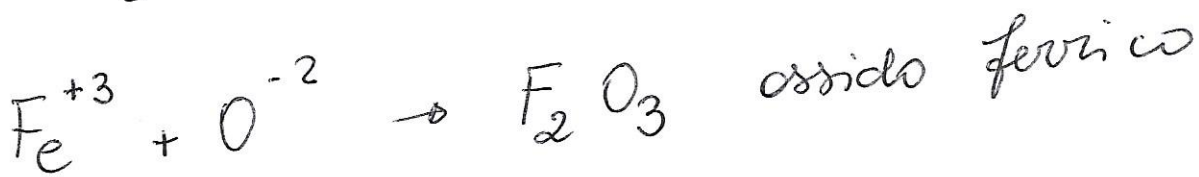
possibili n.o. di Cl : 1, 3, 5, 7
 ipo oso ico per
 oso ↑ ico

Devo usare il n.o. 5

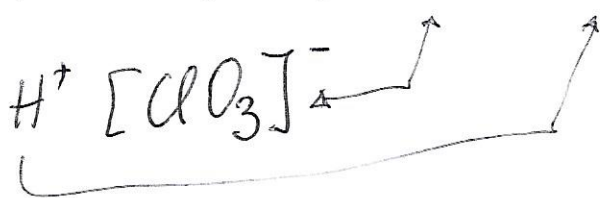
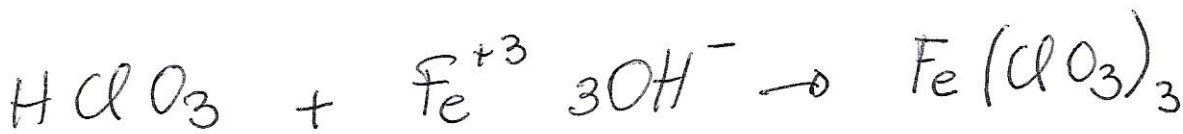


* Per la base parto dal ferro con il suo max n.o di oss.

Fe n.o. 2, (3) quindi 3



SALE = ACIDO + BASE



per bilanciare le 3 cariche positive di Fe^{+3} servono 3 gruppi $[\text{ClO}_3]^-$